

Fig.1

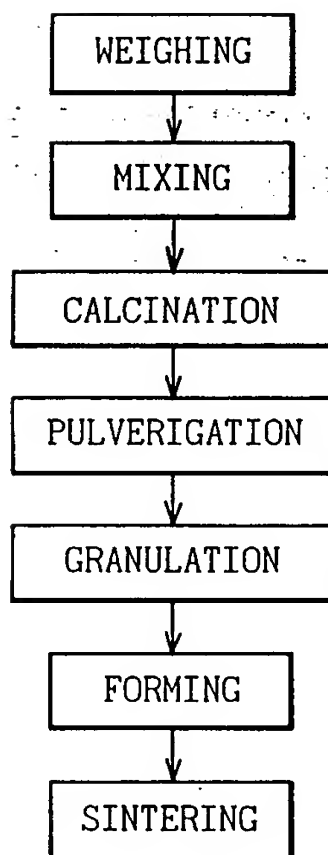


Fig.2

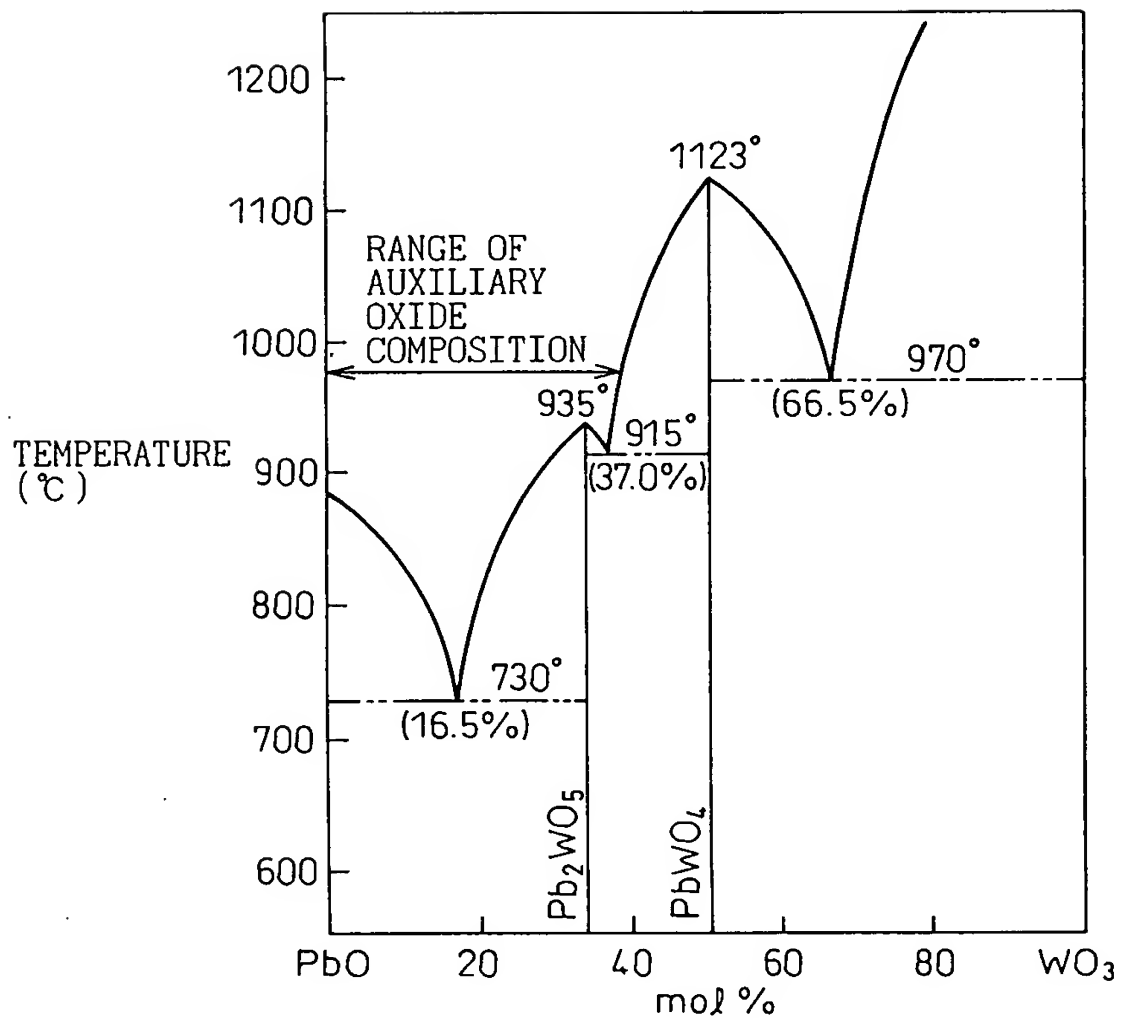


Fig.3

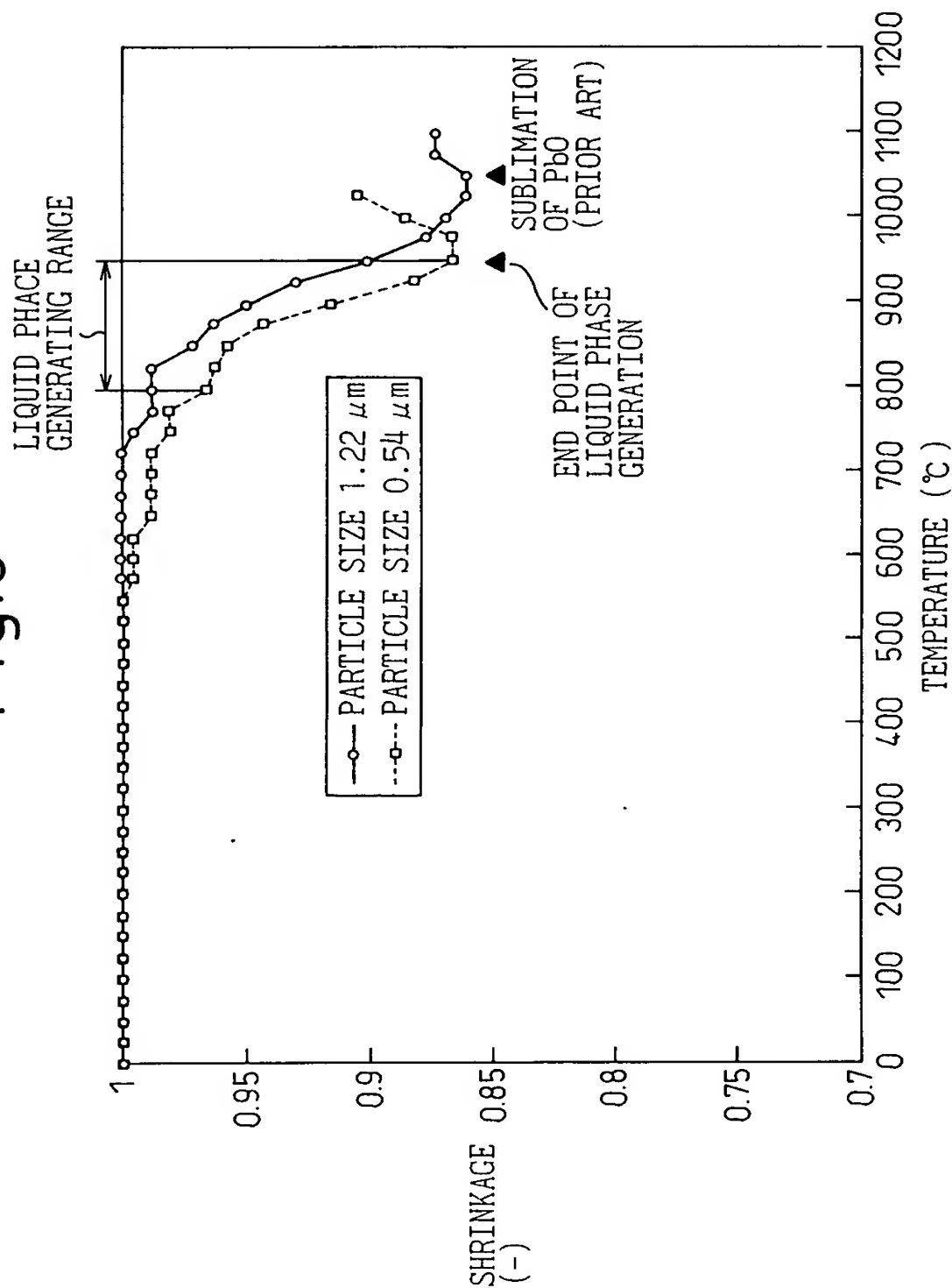


Fig. 4(a)

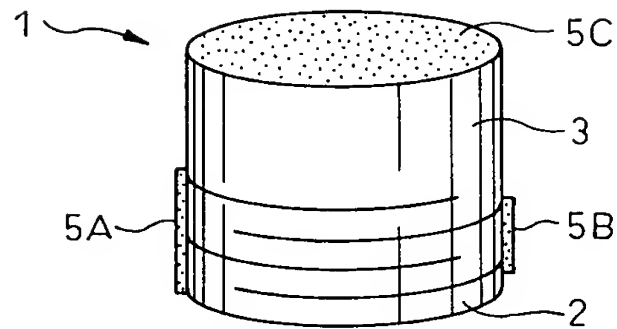


Fig. 4(b)

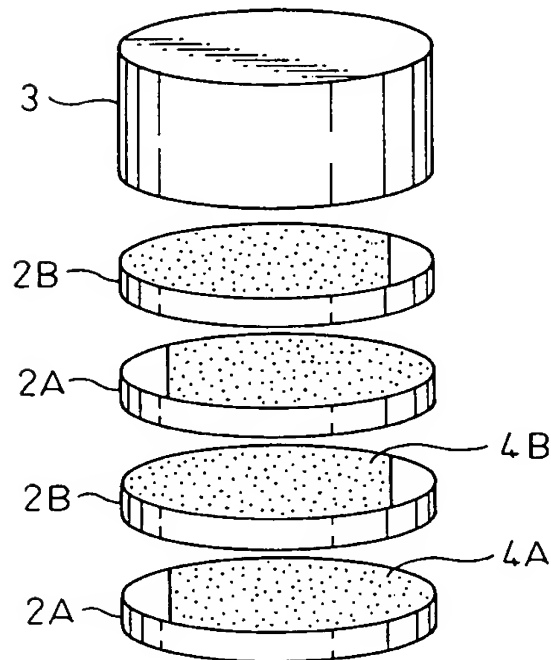


Fig.5(a)

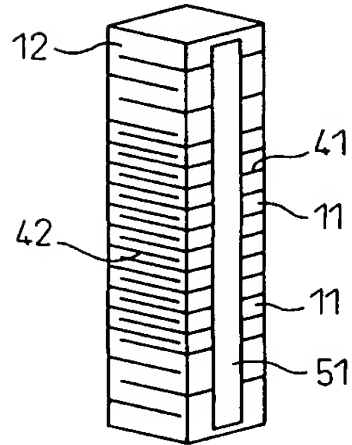


Fig.5(c)

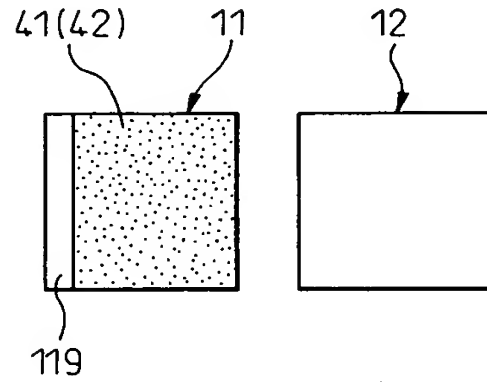


Fig.5(b)

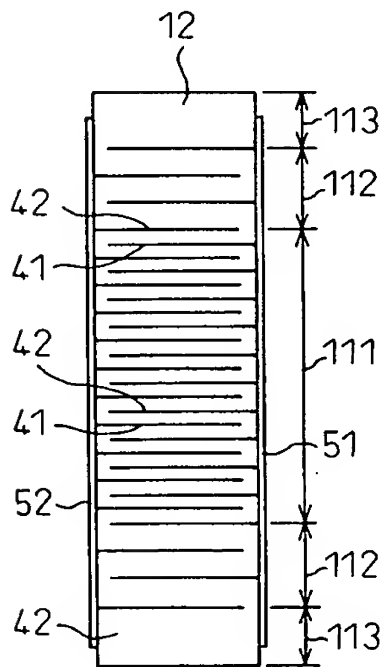


Fig.5(d)

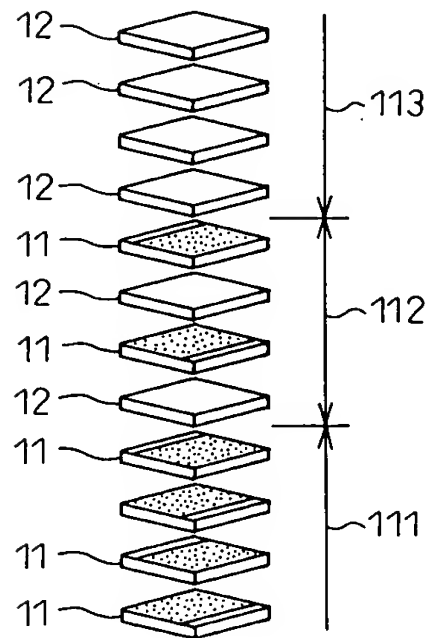


Fig.6(a)

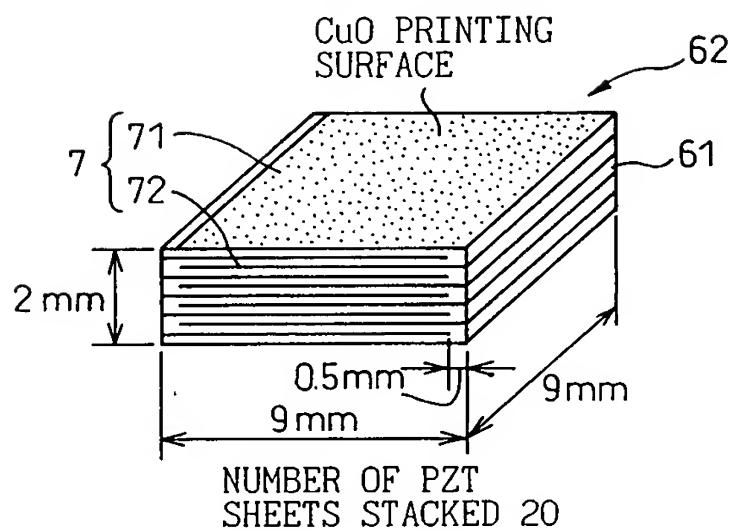
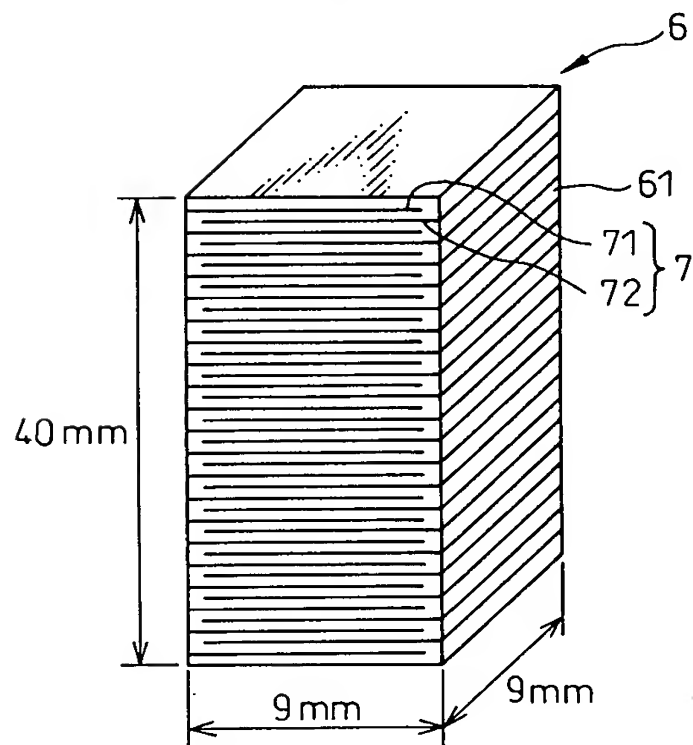


Fig.6(b)



7/12

Fig.7(a)

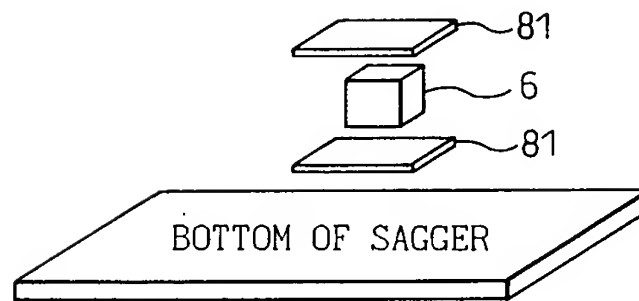


Fig.7(b)

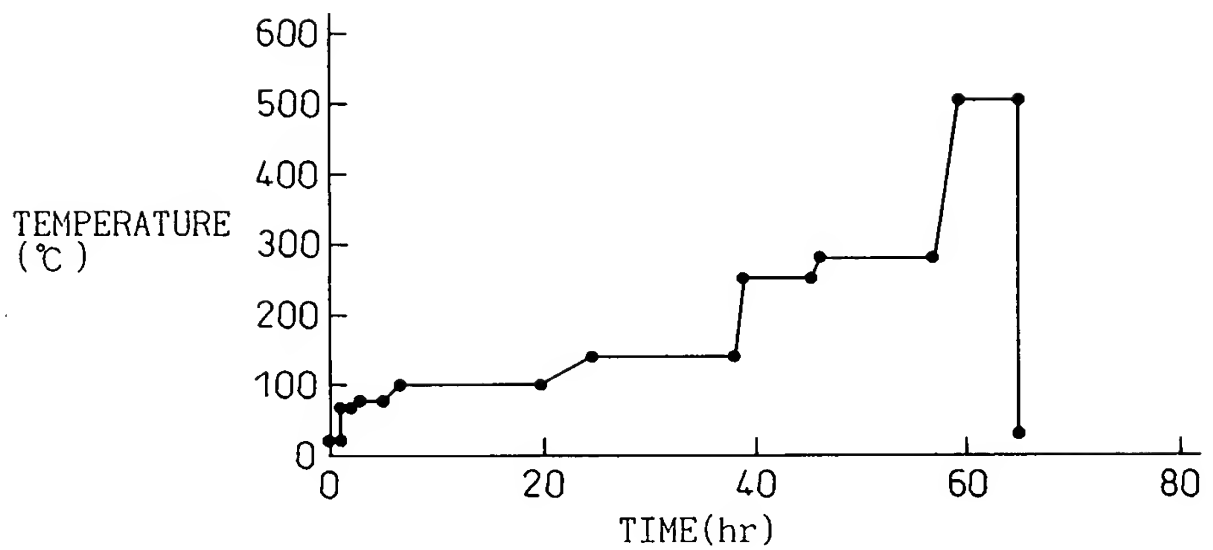


Fig.8

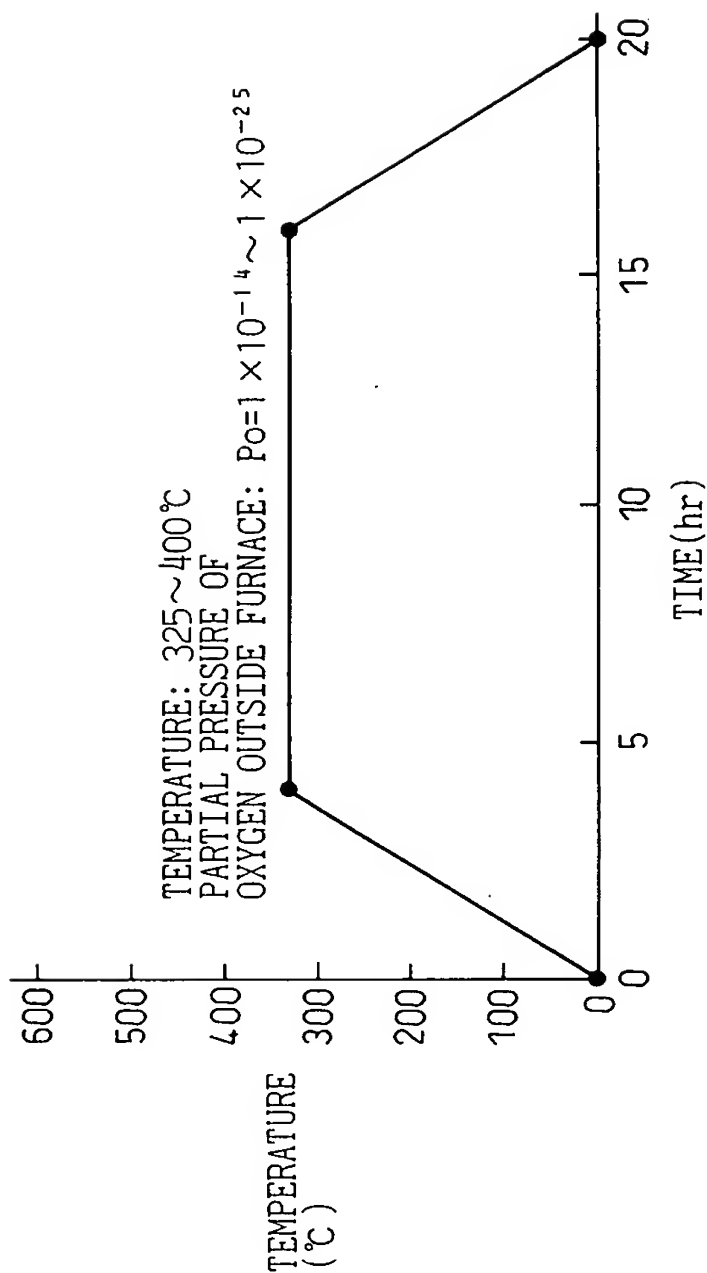


Fig.9(a)

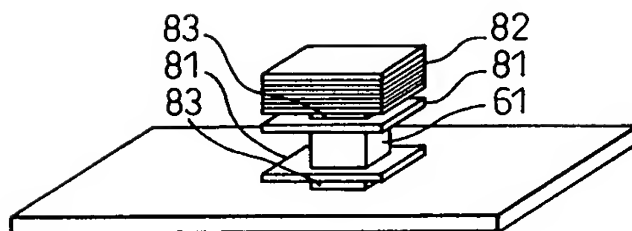


Fig.9(b)

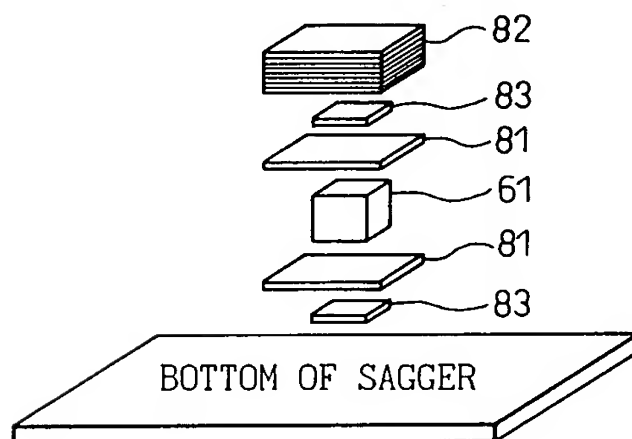


Fig.9(c)

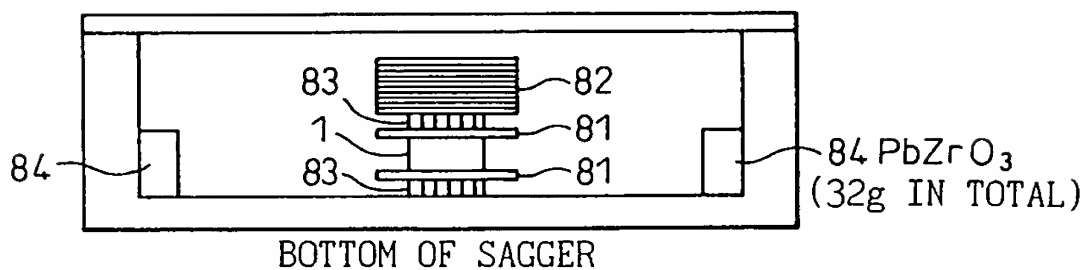
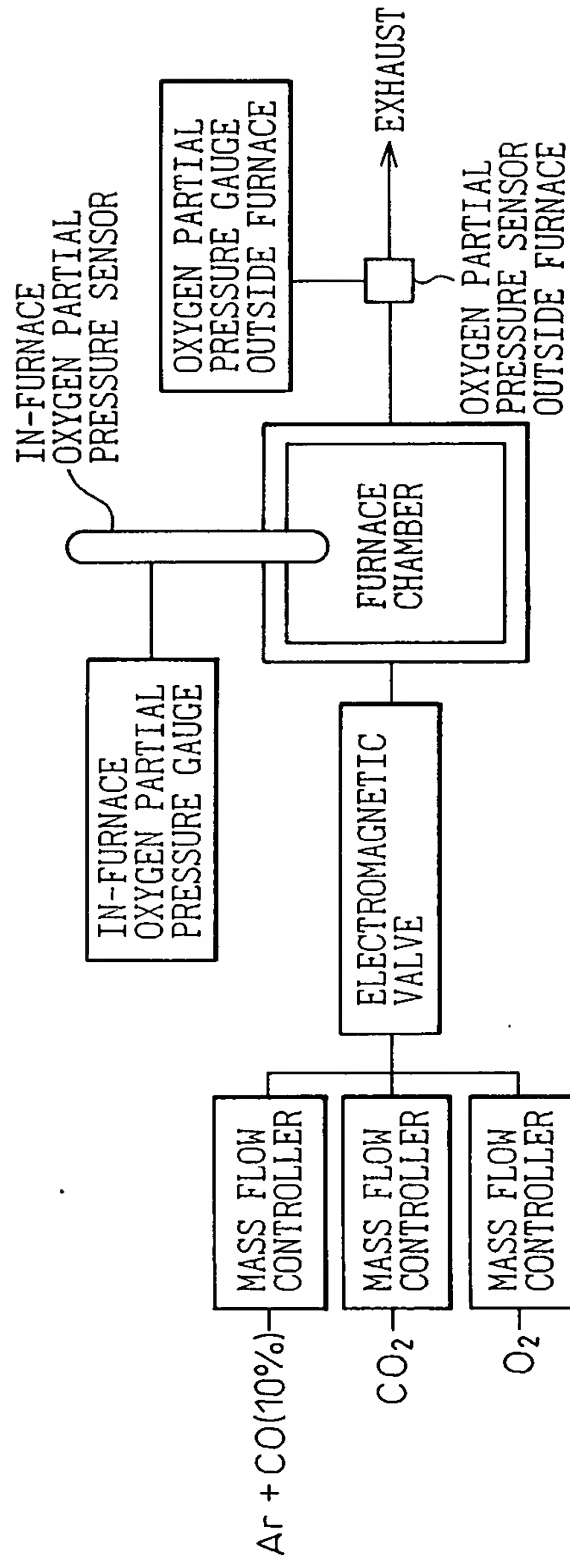


Fig.10



Graph showing Temperature ($^{\circ}\text{C}$) and Partial Pressure of Oxygen (10^x atm) versus Time (hr) for the oxidation of iron at 1000°C .

The graph displays two curves over a 16-hour period:

- Temperature ($^{\circ}\text{C}$):** The temperature is constant at 1000°C throughout the entire duration from 0 to 16 hours.
- Partial Pressure of Oxygen (10^x atm):** The partial pressure starts at 0, rises to 1.0 at 2 hours, and then remains constant at 1.0 until 16 hours.

TEMPERATURE (°C)

PARTIAL PRESSURE OF OXYGEN (10^xatm)

TIME (hr)

PARTIAL PRESSURE OF OXYGEN OUTSIDE FURNACE

PARTIAL PRESSURE OF OXYGEN IN FURNACE

TEMPERATURE

12/12

Fig.11(c)

